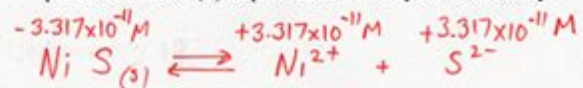


10. The molar solubility of nickel (II) sulphide is  $3.317 \times 10^{-11}$  M. Calculate the value of the solubility product for nickel (II) sulphide. Show all of your work clearly.



$$K_{sp} = [\text{Ni}^{2+}][\text{S}^{2-}]$$

$$K_{sp} = (3.317 \times 10^{-11})^2 = \boxed{1.100 \times 10^{-21}}$$

11. A solution of potassium chloride is titrated with 0.200 M silver nitrate solution. The following data table was obtained:

|                                                | Trial 1 | Trial 2 | Trial 3 |
|------------------------------------------------|---------|---------|---------|
| Initial AgNO <sub>3</sub> burette reading (mL) | 0.00    | 5.26    | 14.63   |
| Final AgNO <sub>3</sub> burette reading (mL)   | 5.26    | 12.19   | 19.87   |
| Volume of KCl titrated                         | 25.0    | 25.0    | 25.0    |

Use the information in the data table to calculate the [Cl<sup>-</sup>] in the KCl solution. Show all of your work clearly.

|                                  | Trial 1 | Trial 2         | Trial 3 |
|----------------------------------|---------|-----------------|---------|
| Vol. AgNO <sub>3</sub> used (mL) | 5.26    | 6.93<br>DISCARD | 5.24    |

$$\text{Best average volume} = \frac{5.26 + 5.24}{2} = 5.25 \text{ mL} \\ = 0.00525 \text{ L AgNO}_3$$

$$\text{moles of AgNO}_3 = 0.200 \text{ M} \times 0.00525 \text{ L} = 0.00105 \text{ mol AgNO}_3 \text{ (Ag}^{\text{+}}\text{)}$$



$$\text{moles KCl} = 0.00105 \text{ mol AgNO}_3 \times \frac{1 \text{ mol KCl}}{1 \text{ mol AgNO}_3} \\ = 0.00105 \text{ mol KCl}$$

$$[\text{Cl}^-] = [\text{KCl}] = \frac{0.00105 \text{ mol}}{0.0250 \text{ L}} = \boxed{0.0420 \text{ M}}$$

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