

DIRECTIONS

- When you have selected your answer, blacken the corresponding space on the answer sheet with a soft, black #2 pencil. Make a heavy, full mark, but no stray marks. If you decide to change an answer, erase the unwanted mark very carefully.
- Make no marks in the test booklet. Do all calculations on scratch paper provided by your examiner.
- There is only one correct answer to each question. Any questions for which more than one response has been blackened will not be counted.
- Your score is based solely on the number of questions you answer correctly. It is to your advantage to answer every question.
- The best strategy is to arrive at your own answer to a question before looking at the choices. Otherwise, you may be misled by plausible, but incorrect, responses.

1. Which gas is **most** soluble in water?
 (A) ammonia (B) hydrogen
 (C) methane (D) nitrogen
2. For which family of elements does the melting point increase with increasing atomic mass?
 1. alkali metals
 2. halogens
 (A) 1 only (B) 2 only
 (C) both 1 and 2 (D) neither 1 nor 2
3. Which pair of substances could be separated by a simple distillation in a school laboratory?
 (A) ammonia and argon (B) salt and water
 (C) oil and vinegar (D) flour and sugar
4. Which compound is **least** soluble in water?
 (A) $\text{Al}(\text{NO}_3)_3$ (B) Na_3PO_4
 (C) K_2SO_4 (D) PbCl_2
5. An object is dried by heating it in an oven. Which procedure would give the most accurate determination of its mass?
 (A) The object is weighed while it is still warm.
 (B) The object is weighed after being cooled on the laboratory bench.
 (C) The object is weighed after being cooled in a desiccator.
 (D) The object is weighed after being cooled by blowing air on it.
6. A sample of gas is to be identified by means of its behavior in the presence of a glowing splint. Which gas will **neither** itself burn **nor** cause the splint to burn?
 (A) oxygen (B) nitrogen
 (C) hydrogen (D) methane
7. A compound prepared as a model of hemoglobin is 4.6% iron by mass. If the compound is known to contain a single iron atom, which of these values will be closest to its molar mass?
 (A) 1000 g mol^{-1} (B) 1200 g mol^{-1}
 (C) 1400 g mol^{-1} (D) 1600 g mol^{-1}
8. How many hydrogen atoms are present in 0.046 g of ethanol, $\text{C}_2\text{H}_5\text{OH}$?
- | | Molar Mass, M |
|---------------------------------|-------------------------|
| $\text{C}_2\text{H}_5\text{OH}$ | 46 g mol^{-1} |
- (A) 6.0×10^{20} (B) 1.2×10^{21}
 (C) 3.0×10^{21} (D) 3.6×10^{21}
9. Which of these hydrated salts contains the greatest percentage of water by mass?
- | Hydrated Salt | Molar Mass |
|---|--------------------------|
| $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ | 244 g mol^{-1} |
| $\text{CaCl}_2 \cdot 6\text{H}_2\text{O}$ | 219 g mol^{-1} |
| $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$ | 250 g mol^{-1} |
| $\text{Ni}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$ | 366 g mol^{-1} |
- (A) $\text{BaCl}_2 \cdot 2\text{H}_2\text{O}$ (B) $\text{CuSO}_4 \cdot 5\text{H}_2\text{O}$
 (C) $\text{CaCl}_2 \cdot 6\text{H}_2\text{O}$ (D) $\text{Ni}(\text{ClO}_4)_2 \cdot 6\text{H}_2\text{O}$
10. Ammonium dichromate, $(\text{NH}_4)_2\text{Cr}_2\text{O}_7$, decomposes when heated to produce N_2 , H_2O , and Cr_2O_3 as represented by this equation.
 $__ (\text{NH}_4)_2\text{Cr}_2\text{O}_7(s) \rightarrow __ \text{N}_2(g) + __ \text{H}_2\text{O}(g) + __ \text{Cr}_2\text{O}_3(s)$
 What is the coefficient for H_2O when this equation is correctly balanced using the smallest set of whole number coefficients?
 (A) 1 (B) 2 (C) 4 (D) 8
11. What volume of 0.15 M HCl can be made from 7.5 mL of concentrated HCl (12 M)?
 (A) 0.060 L (B) 0.60 L
 (C) 6.0 L (D) $6.0 \times 10^2 \text{ L}$

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