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## CHAPTER 11 REVIEW

*Molecular Composition of Gases*

## MIXED REVIEW

**SHORT ANSWER** Answer the following questions in the space provided.

- \_\_\_\_\_ The average speed of a gas molecule is most directly related to the \_\_\_\_\_.
  - polarity of the molecule
  - pressure of the gas
  - temperature of the gas
  - number of moles in the sample
- \_\_\_\_\_ Which of the following statements best represents Graham's law for two gases at the same temperature?
  - The molar masses and the velocities are inversely proportional.
  - The molar masses and the velocities are directly proportional.
  - The molar masses and the square roots of the velocities are directly proportional.
  - The molar masses and the squares of the velocities are inversely proportional.
- Water is synthesized from its elements according to the following equation:
 
$$2\text{H}_2(\text{g}) + \text{O}_2(\text{g}) \rightarrow 2\text{H}_2\text{O}(\text{l})$$
  - Which is greater, the volume of  $\text{H}_2$  consumed or the volume of  $\text{O}_2$  consumed?  
\_\_\_\_\_
  - Which is greater, the mass of  $\text{H}_2$  consumed or the mass of  $\text{O}_2$  consumed?  
\_\_\_\_\_
- Consider a 2.24 L sample of an ideal gas at STP.
  - How many moles are present in the sample, or is more data needed?  
\_\_\_\_\_
  - How many molecules are present in the sample, or is more data needed?  
\_\_\_\_\_
  - What is the mass of the sample, or is more data needed?  
\_\_\_\_\_

**PROBLEMS** Write the answer on the line to the left. Show all your work in the space provided.

- \_\_\_\_\_ Calculate the volume of 3.5 mol of an ideal gas at  $24^\circ\text{C}$  and 760 torr.

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