

Name: _____
Hour: _____ Date: _____

Chemistry: Colligative Properties – Freezing Point Depression

1. How much will the freezing point be lowered if enough sugar is dissolved in water to make a 0.50 molal solution?
2. What is the freezing point of a solution of a nonelectrolyte dissolved in water if the concentration of the solution is 0.24 *m*?
3. What is the freezing point of a solution that contains 68.4 g of sucrose, C₁₂H₂₂O₁₁, dissolved in 1.00x10² g of water?
4. Suppose that 98.0 g of a nonelectrolyte is dissolved in 1.00 kg of water. The freezing point of this solution is found to be -0.465°C. What is the molecular mass of the solute?
5. A researcher places 53.2 g of an unknown nonelectrolyte in 505 g naphthalene. The nonelectrolyte lowers naphthalene's freezing point by 8.8°C. What is the molar mass of the unknown substance?

ANSWERS: 1. 0.93°C 2. -0.446°C 3. -3.72°C 4. 392 amu 5. 81.4 g/mol

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