

Name: Courtney Wallace
 Honors Chemistry Quiz: Gas Laws

45 Pts.

- 1.) Write the correct answer in the space left. (15 pts.)
- 18245.6 What is the pressure in Pascal's if an area of 0.285m^2 has a force of 5200N exerted on it?
 - Barometer Term for device which measures air pressure.
 - -273°C At what Celsius temperature do ideal gases have no volume? $\frac{712}{1} \times \frac{1}{760}$
 - $.937\text{atm}$ At high elevations the air pressure is 712 torr. What is this pressure in atmospheres?
 - Charles's The relationship between the volume and temperature of a gas was studied by ___.
 - Boyle's The relationship between volume and pressure of a gas was studied by ___.
 - doubled To decrease the volume of a gas from 10L to 5L the pressure must be (doubled/halved).
 - -86.5°C What Celsius temperature is needed to halve the volume of a gas initially at 100°C ? $\frac{12.0\text{L}}{673} = \frac{x}{373}$
 - Same Warm gas molecules are (larger / smaller / the same size) than/as cool gas molecules.
 - 5.12L 12.0L of gas is held at 600°C . If the temperature is decreased to 100°C what is the new volume?
 - 4mol H_2 Which gas exerts the greatest pressure at 20°C and 1L ? (4mol H_2 , 2mol He , 0.5mol CH_4) $\frac{4}{22.4} > \frac{2}{22.4} > \frac{0.5}{22.4}$
 - 1mol O_2 Which will have the greatest change in volume when heated? 1.0L Fe , $1.0\text{L H}_2\text{O}$, 1.0L O_2
 - 22.4L What volume will 1.00 mole of an ideal gas occupy at 0°C and 1atm ?
 - Inversely At constant pressure and volume the temperature of a gas is (inversely / directly) related to the number of moles.
 - directly At constant moles and volume the pressure of a gas is (inversely / directly) related to its absolute temperature.

- 2.) By what factor will an ideal gases volume change if it's pressure changes from 1182 torr to 730 torr while it's temperature changes from 25°C to 100°C ? (Use a written statement such as "the volume will (increase/decrease) by a factor of..."). (4 pts.)

$$\frac{22.4\text{L} \cdot 1182\text{torr}}{298} = \frac{x \cdot 730\text{torr}}{373} \quad \left(\frac{1182}{730}\right)\left(\frac{373}{298}\right) = 2.03$$

$$9875846.4 = 217340x$$

- 3.) Which is/are characteristics of an ideal gas? Circle letter of all that agree. (3 pts.)
- Ideal gases are point-masses ✓
 - Ideal gases behave best at low pressure and high temperatures.
 - Collisions among gas molecules are elastic.
 - At absolute zero ideal gases have no volume.
 - Gases behave most ideally when condensed to a liquid or solid.
 - 1 mole of an ideal gas occupies 22.4L .
- 4.) 518 grams of N_2O gas is held in a tank at 20.0°C and 4.20atm . What mass of P_2O_5 gas would there be in this same tank at these same conditions? (3 pts.)

$$\frac{518}{1} \cdot \frac{1\text{mole}}{44} = \frac{11.77\text{ moles}}{1} \times \frac{142}{1\text{mole}} = 1670\text{g}$$

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