

Multiple Choice Questions

There is only one correct answer to each question.

1. Which of the following square planar complex ions can have cis-trans isomers?
A. $[\text{Pt}(\text{NH}_3)_4]^{2+}$
B. $[\text{Ni}(\text{NH}_3)_4]^{2+}$
C. $[\text{Pt}(\text{NH}_3)_2(\text{SCN})_2]$
D. $[\text{Pt}(\text{NH}_3)(\text{SCN})_3]^-$
E. $[\text{Ni}(\text{NH}_3)_3(\text{SCN})]^+$
2. A complex with the composition $[\text{MA}_2\text{B}_2]\text{X}_2$ is found to have no geometrical isomers. Both A & B are monodentate ligands. The structure of the complex is:
A. linear
B. square planar
C. tetrahedral
D. octahedral
3. Apply the VSEPR model to XeF_4 . Which of the following molecular shapes is consistent with the model?
A. Square planar
B. K-shaped
C. Tetrahedral
D. Square pyramidal
E. Tetragonal
4. In the complex ion ML_6^{n+} , M^{n+} has seven d electrons and L is a strong field ligand. According to crystal field theory, the magnetic properties of the complex ion correspond to how many unpaired electrons?
A. 0
B. 1
C. 2
D. 3
E. 5

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